

## Lots of Ionic Naming Practice Problems

Name the following ionic compounds:

- 1) NaBr \_\_\_\_\_
- 2) Sc(OH)<sub>3</sub> \_\_\_\_\_
- 3) V<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub> \_\_\_\_\_
- 4) NH<sub>4</sub>F \_\_\_\_\_
- 5) CaCO<sub>3</sub> \_\_\_\_\_
- 6) NiPO<sub>4</sub> \_\_\_\_\_
- 7) Li<sub>2</sub>SO<sub>3</sub> \_\_\_\_\_
- 8) Zn<sub>3</sub>P<sub>2</sub> \_\_\_\_\_
- 9) Sr(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>2</sub> \_\_\_\_\_
- 10) Cu<sub>2</sub>O \_\_\_\_\_
- 11) Ag<sub>3</sub>PO<sub>4</sub> \_\_\_\_\_
- 12) YClO<sub>3</sub> \_\_\_\_\_
- 13) SnS<sub>2</sub> \_\_\_\_\_
- 14) Ti(CN)<sub>4</sub> \_\_\_\_\_
- 15) KMnO<sub>4</sub> \_\_\_\_\_
- 16) Pb<sub>3</sub>N<sub>2</sub> \_\_\_\_\_
- 17) CoCO<sub>3</sub> \_\_\_\_\_
- 18) CdSO<sub>3</sub> \_\_\_\_\_
- 19) Cu(NO<sub>2</sub>)<sub>2</sub> \_\_\_\_\_
- 20) Fe(HCO<sub>3</sub>)<sub>2</sub> \_\_\_\_\_

Write the formulas for the following ionic compounds:

- 21) lithium acetate \_\_\_\_\_
- 22) iron (II) phosphate \_\_\_\_\_
- 23) titanium (II) selenide \_\_\_\_\_
- 24) calcium bromide \_\_\_\_\_
- 25) gallium chloride \_\_\_\_\_
- 26) sodium hydride \_\_\_\_\_
- 27) beryllium hydroxide \_\_\_\_\_
- 28) zinc carbonate \_\_\_\_\_
- 29) manganese (VII) arsenide \_\_\_\_\_
- 30) copper (II) chlorate \_\_\_\_\_
- 31) cobalt (III) chromate \_\_\_\_\_
- 32) ammonium oxide \_\_\_\_\_
- 33) potassium hydroxide \_\_\_\_\_
- 34) lead (IV) sulfate \_\_\_\_\_
- 35) silver cyanide \_\_\_\_\_
- 36) vanadium (V) nitride \_\_\_\_\_
- 37) strontium acetate \_\_\_\_\_
- 38) molybdenum sulfate \_\_\_\_\_
- 39) platinum (II) sulfide \_\_\_\_\_
- 40) ammonium sulfate \_\_\_\_\_